

NCERT/CBSE CHEMISTRY CLASS 11 textbook

<http://www.TutorBreeze.com>

Contact for Online Tutoring in Physics, Math, Chemistry,

English, French, Accounts, MBA

Solutions/Answers to NCERT/CBSE CHEMISTRY Class 11(Class XI)textbook

CHAPTER TEN

The s-Block Elements

10.6 Compare the alkali metals and alkaline earth metals with respect to (i) ionization enthalpy (ii) basicity of oxides and (iii) solubility of hydroxides.

- (i) Ionization enthalpy: - The ionization enthalpy of alkaline earth metal is higher than that of alkali metals because the alkaline earth metal has higher nuclear charge and smaller atomic size as compared to alkali metals.
- (ii) Basicity of oxides: - The electropositive character of alkali metals is more than that of alkaline earth metals. Therefore the oxides of alkali metals are more basic than the oxides of alkaline earth metals.
- (iii) Solubility of hydroxides: - The hydroxides of alkali metals are more than that of the hydroxides of alkaline earth metals. This is because the hydroxides of alkali metals have higher ionic charge and smaller size of the metals as compared to hydroxides of alkaline earth metals. Therefore the lattice enthalpies of hydroxides of alkali is more than that of hydroxides of alkaline earth metals.

©TutorBreeze.com

Please do not copy the answer given here

[Write to us for help in understanding the solution](#)