

NCERT/CBSE CHEMISTRY CLASS 11 textbook

<http://www.TutorBreeze.com>

Contact for Online Tutoring in Physics, Math,

Chemistry, English, Accounts, MBA

Solutions/Answers to NCERT/CBSE CHEMISTRY Class 11(Class XI)textbook

CHAPTER Four

CHEMICAL BONDING AND MOLECULAR STRUCTURE

4.18 Explain with the help of suitable example polar covalent bond.

Solution

Polar covalent bonds are the covalent bonds in which the atoms forming the bond have unequal electronegativities. As a result the electron pair is attracted towards the atom with higher electronegativity. For example in HF, the electron pair is attracted towards F because of it has higher electronegativity than H. HF can be written as $H^{+\delta} F^{-\delta}$, where δ is a small charge.

©TutorBreeze.com

Please do not copy the answer given here

[Write to us for help in understanding the solution](#)