

NCERT/CBSE CHEMISTRY CLASS 11 textbook

<http://www.TutorBreeze.com>

Contact for Online Tutoring in Physics, Math, Chemistry,

English, Accounts, MBA

Solutions/Answers to NCERT/CBSE CHEMISTRY Class 11(Class XI)textbook

CHAPTER SIX THERMODYNAMICS

- 6.7 In a process, 701 J of heat is absorbed by a system and 394 J of work is done by the system. What is the change in internal energy for the process?

1. Solution:

Change in internal energy, $\Delta E = q + w = 701 - 394 = 307$ J

©TutorBreeze.com

Please do not copy the answer given here

[Write to us for help in understanding the solution](#)