

NCERT/CBSE CHEMISTRY CLASS 11 textbook

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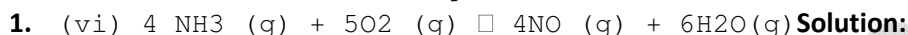
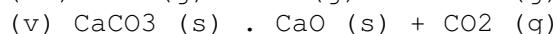
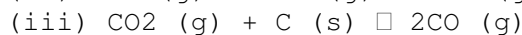
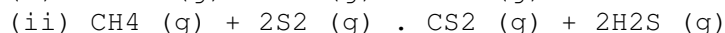
Solutions/Answers to NCERT/CBSE CHEMISTRY Class 11(Class XI)textbook

CHAPTER SEVEN

EQUILIBRIUM

7.26 Which of the following reactions will get affected by increasing the pressure?

Also, mention whether change will cause the reaction to go into forward or backward direction.



(i) $\Delta n_g = 2 - 1 = 1$. Hence reaction shifts in backward direction on increasing the pressure.

(ii) $\Delta n_g = 0$. Hence there is no effect on increasing the pressure.

(iii) $\Delta n_g = 2 - 1 = 1$. Hence reaction shifts in backward direction on increasing the pressure.

(iv) $\Delta n_g = 1 - 3 = -2$. Hence reaction shifts in forward direction on increasing the pressure.

(v) $\Delta n_g = 1 - 0 = 1$. Hence reaction shifts in backward direction on increasing the pressure.

(vi) $\Delta n_g = 10 - 9 = 1$. Hence reaction shifts in backward direction on increasing the pressure.