

NCERT/CBSE CHEMISTRY CLASS 11 textbook

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CHAPTER SEVEN

EQUILIBRIUM

7.6 For the following equilibrium, $K_c = 6.3 \times 10^{14}$ at 1000 K



1.Solution:

$K_c = 6.3 \times 10^{14}$ (given) at 1000 K for forward reaction.

For the reverse reaction, $K_c' = \frac{1}{K_c} = \frac{1}{6.3 \times 10^{14}} = 1.59 \times 10^{-15}$