

# NCERT/CBSE CHEMISTRY CLASS 11 textbook

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Solutions/Answers to NCERT/CBSE CHEMISTRY Class 11 (Class XI) textbook

## CHAPTER EIGHT

### REDOX REACTIONS

8.1 Assign oxidation number to the underlined elements in each of the following

species:

(a)  $\text{NaH}_2\text{PO}_4$  :- Oxidation number of P in  $\text{NaH}_2\text{PO}_4$  is  $1(+1)+2(+1)+1(x)+4(-2)=0$   
 $= X = +5$

(b)  $\text{NaHSO}_4$  :- Oxidation number of S in  $\text{NaHSO}_4$  is  $1(+1)+1(+1)+x+4(-2)=0$   
 $= x = +6$

(c)  $\text{H}_4\text{P}_2\text{O}_7$  :- Oxidation number of P in  $\text{H}_4\text{P}_2\text{O}_7$  is  $4(+1)+2(x)+7(-2)=0$   
 $= x = +5$

(d)  $\text{K}_2\text{MnO}_4$  :- Oxidation number of Mn in  $\text{K}_2\text{MnO}_4$  is  $2(+1)+1(x)+4(-2)=0$   
 $= X = +7$

(e)  $\text{CaO}_2$  :- Oxidation number of O in  $\text{CaO}_2$  is  $+2+2x=0$   
 $= x = -1$

(f)  $\text{NaBH}_4$  :- Oxidation number of B in  $\text{NaBH}_4$  is  $1(+1) + x + 4(-1) = 0$   
 $= x = +3$

(g)  $\text{H}_2\text{S}_2\text{O}_7$  :- Oxidation number of S in  $\text{H}_2\text{S}_2\text{O}_7$  is  $2(+1)+2(x)+7(-2)=0$   
 $= x = +6$

(h)  $\text{KAl(SO}_4)_2 \cdot 12\text{H}_2\text{O}$  :- Oxidation number of S in  $\text{KAl(SO}_4)_2 \cdot 12\text{H}_2\text{O}$  is  
 $+1+3+2x-16=0 = x = +6$