

## NCERT/CBSE PHYSICS CLASS 11 textbook

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Exercise CHAPTER TWO

SOLUTIONS

EXERCISES

2.23

Suggest the most important type of intermolecular attractive interaction in the following pairs.

(i) n-hexane and n-octane

Both are non polar , so the intermolecular forces in them will be London dispersion force.

(ii) I<sub>2</sub> and CCl<sub>4</sub>

Both are non polar , so the intermolecular forces in them will be London dispersion force

(iii) NaClO<sub>4</sub> and water

NaClO<sub>4</sub> gives sodium ions and chlorate ions in solution. Water is a polar molecule, so the intermolecular forces in them will be ion-Dipole interactions.

(iv) methanol and acetone

Both are polar , so the intermolecular forces in them will be Dipole -Dipole interactions

(v) acetonitrile (CH<sub>3</sub>CN) and acetone (C<sub>3</sub>H<sub>6</sub>O).

Both are polar , so the intermolecular forces in them will be Dipole -Dipole interactions