

NCERT/CBSE PHYSICS CLASS 11 textbook

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CHAPTER THREE

ELECTROCHEMISTRY

EXERCISES

3.9

The resistance of a conductivity cell containing 0.001M KCl solution at 298 K is 1500 Ω . What is the cell constant if conductivity of 0.001M KCl solution at 298 K is $0.146 \times 10^{-3} \text{ S cm}^{-1}$.

$$\begin{aligned}\text{Cell Constant} &= \text{Conductivity} \times \text{Resistance} \\ &= 0.146 \times 10^{-3} \times 1500 \\ &= 0.219 \text{ cm}^{-1}\end{aligned}$$

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