

NCERT/CBSE CHEMISTRY CLASS 12 textbook

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CHAPTER SEVEN

The p-Block Elements

7.18

Why is dioxygen a gas but sulphur a solid? .

Di oxygen is a gas because

- (1) Oxygen is small in size and highly electronegative.
- (2) It forms p π - p π multiple bonds.
- (3) The oxygen molecules are held to each other by weak vanderwaals forces of attraction.

Sulphur is a solid because

- (1) Sulphur is bigger in size and less electronegative.
- (2) It does not form p π - p π multiple bonds, but forms S-S single bond which is stronger .
- (3) Sulphur has the tendency to undergo catenation and therefore forms octa atomic molecule .
- (4) The bigger size , strong force of attraction are much stronger in Sulphur.