

# NCERT/CBSE CHEMISTRY CLASS 12 textbook

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## CHAPTER EIGHT

### The d- and f-Block Elements

#### 8.11

Explain giving reasons:

(i) Transition metals and many of their compounds show paramagnetic behaviour.

This is due to the presence of unpaired electrons.

(ii) The enthalpies of atomisation of the transition metals are high.

This is due to strong metallic bonds and additional covalent bond due to the presence of unpaired electrons.

(iii) The transition metals generally form coloured compounds.

This is due to the presence of unpaired d-electrons. They undergo d-d transition by absorbing the light from the visible region. And then radiate the coloured light which is complimentary.

(iv) Transition metals and their many compounds act as good catalyst.

This is due to the variable oxidation states exhibited by the transition elements